We have been able to describe the properties of solids rather completely. The properties of liquids are more difficult to predict. At low energy levels, liquids are "solid like." Lattice vibrations predominate over translation of the particles from one region to another within the liquid. Thus the energy and specific heats of atomic and molecular liquids are not too different from those of solids just discussed. A major difference is that we must add an entropy component to account for the fact that liquid particles are nonlocalized indistinguishable particles. One difficulty in analyzing liquid properties is determining just how much of this "communal entropy" to include in evaluating the entropy of liquids. (The full communal entropy is the difference between the Maxwell-Boltzmann and Boltzmann limit entropies, $s_{comm} = k \ln N!$.) As the energy level of the liquid increases, a greater percentage of the particles are able to escape to new positions within the liquid lattice. In these instances a fixed neighbor, simple harmonic vibration model for the lattice energy is increasingly complicated. It may even be necessary to abandon the independent particle assumption. We leave quantitative determination of liquid properties to more advanced treatments.